

« » 2016

1.  $1s^2 2s^2 2p^6 3s^2 3p^3$ .  $^{31}_{15}\text{P}$

2.  $-\text{Cr}(\text{OH})_3, \text{Zn}(\text{OH})_2, \text{Al}(\text{OH})_3$

3. (25,8%), (7,5%), Dep (4,7%), (3,4%), (49,4%), (2,6%) ін.

4. Б  $^{204}_{82}\text{Pb} + ^4_2\text{He}$  :  $^{208}_{84}\text{Po}$

5. З  $2, 2, \text{N}_2$

6. р.  $\text{H}_2^{+1}\text{C}^{+4}\text{O}_3^{-2}$

7.  $:\text{K}_2^{+1}\text{Cr}_2^{+6}\text{O}_7^{-2}$

8.  $(\text{ }_6\text{ }_{10}\text{ }_5)n$

9.  $\text{HCl} - , \text{H}_2\text{SiO}_3 - , \text{SO}_4 - , \text{H}_3\text{PO}_4 - , \text{HNO}_2 -$

10.  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O} - \text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$

11.  $-\text{N}_2\text{O}$

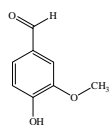
12.  $3 -$

13.  $:\text{C}_n\text{H}_{2n+2}$

14.  $-\text{CH}_2\text{OH}$   
 $|\text{CH}_2\text{OH}$

15.  $4 + \text{HNO}_3( ) 3 - \text{NO}_2 + \text{H}_2$

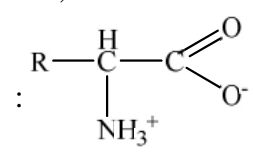
16.  $4 + \text{HNO}_3( ) 3 - \text{NO}_2 + \text{H}_2$



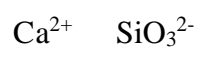
19.

20.

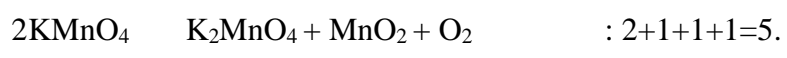
21.



22.



23.



24.

25.

альдегідні групи ..... оксонова кислота .....

26.

на газу та газом визначається за формулу:  $D_{за газом B(A)} = \frac{M_r(A)}{M_r(B)} = \frac{M(A)}{M(B)} = \frac{\rho(A)}{\rho(B)}$  ;  $D_{CO_2(N_2O)} = \frac{44}{44} = 1.$

27.

6 5

28.

3

29.

: ( )<sub>2</sub> + 2      3 + 2

30.

100      100 · 0,1 = 10      : 100  
- 10 = 90      10 : 90 = 1 : 9.

31.

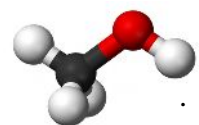
о відносна  
:  $\frac{56}{x} = \frac{44,8}{22,4} \Rightarrow x = 28.$

32.

33. А

R-COO

34.



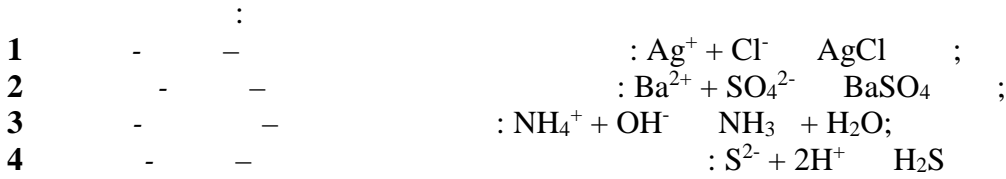
35.

36.

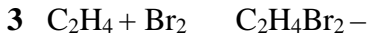
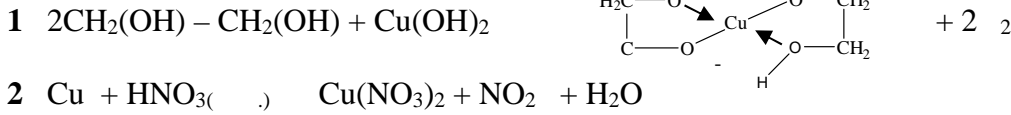
37.

38.

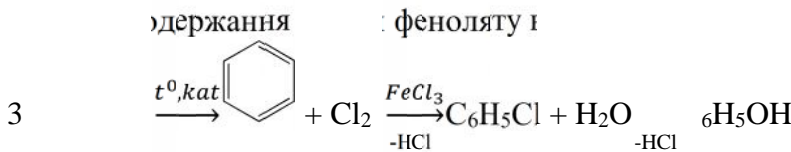
39.



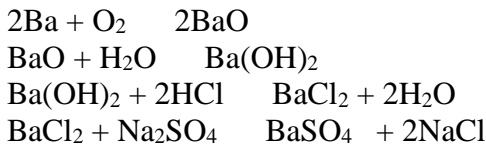
40.



41.



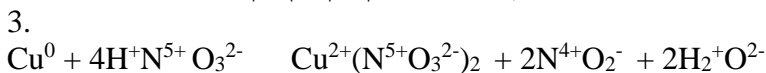
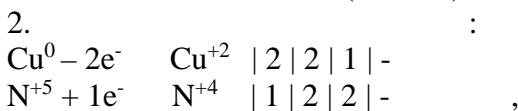
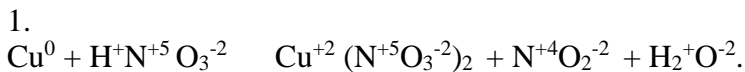
42.



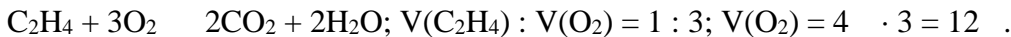
43. 2

$M_r(MeO) = 40 \Rightarrow A_r(Me) = 40 - 16 = 24 \Rightarrow Me - Mg$  2

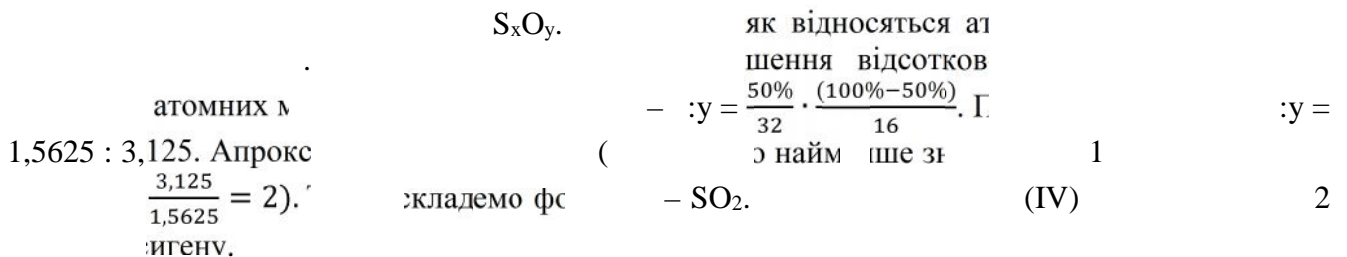
44. 5



45. 12



46.I. 2



46.II. 5



$m_{розчину} = 500 \cdot 1,8 = 900g$



$$W = \frac{m_{\text{речовини}}}{m_{\text{розчину}}} \cdot 100\%$$

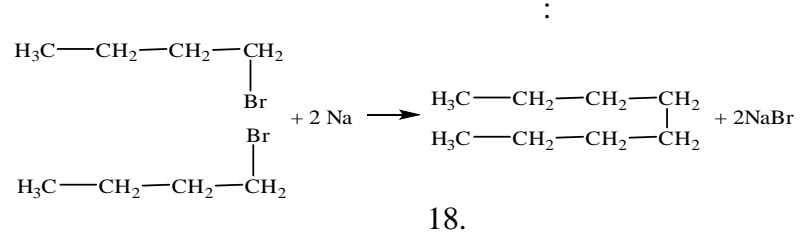
$$W(H_2SO_4) = \frac{45}{900} \cdot 100\% = 5\%$$

**47.I. 58**



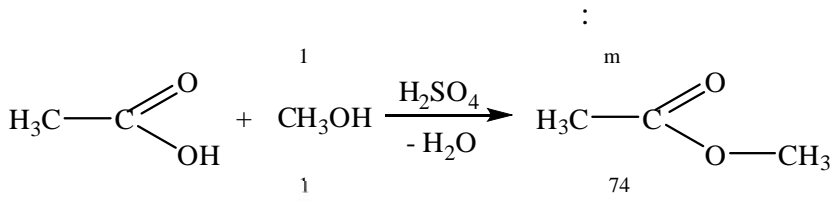
1.  $n(Na) = 46 : 23 / = 2$  ;
2.  $n(C_2H_5Br) = 272,5 \cdot 0,8 : 109 / = 2$  ;
3.  $m(C_4H_{10}) = 2 \cdot 58 / : 2 = 58$  .

**47.II. 18**



**48. 75**

1. C



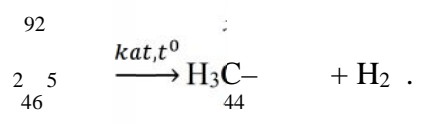
2. , , ікої д'звол.
- $m = 74$  .
3. :  $W = \frac{55,5}{74} \cdot 100\% = 75\%$ .

**49. 6**

1. , , b, c, d. формула ма сполуці, розд - :  $b:c:d = \frac{39}{39} : \frac{12}{12} : \frac{48}{16} : \frac{1}{1}$  .
2. :  $b:c:d = 1:1:3:1$  .
3. « » , си -1 ц , Сума дч 6 (1+1+1+3=6).

**50. 80**

1.  $W_{\text{вих.}} = \frac{m_i}{m} \cdot 100\%$  , ю реакц . Тому, те



2. і кроком о ма теорі .  
 $x = \frac{92 \cdot 44}{46} = 88\text{г.}$
3. :  $W_{\text{вих.}} = \frac{70,4}{88} \cdot 100\% = 80\%$ .